



Student's name: _____

Year: 5th "B"

TRABAJO PRACTICO: "La Tabla Periódica"

1. Read the first paragraph of the text and complete with the correct Past Simple form of the verbs in brackets.

The Periodic Table

As more elements ¹_____ (be) discovered, it ²_____ (become) necessary to organize them into some type of classification system. By the late 1800s, scientists ³_____ (recognize) that certain elements ⁴_____ (look) alike and ⁵_____ (behave) much the same way. In 1869, a Russian chemist, Dmitri Mendeleev, ⁶_____ (arrange) the 60 elements known at that time into groups with similar properties and ⁷_____ (place) them in order of increasing atomic masses. Today, this arrangement of 118 elements is known as the **periodic table**.

2. Choose the correct form of the Simple Past.

- a. What did scientists **recognize / recognized**?
- b. When **was / were** the periodic table arranged?
- c. In 1879 Dmitri Mendeleev **not placed / didn't place** the elements.
- d. Dmitri Mendeleev **didn't organized / didn't organize** the elements with different properties.
- e. **Did it / It did** become necessary to organize the elements?

3. Read the second part of the text and write T (True) or F (False).

Periods and Groups

Each horizontal row in the periodic table is a **period**. The periods are counted from the top of the table as Periods 1 to 7. The **first** period contains two elements: hydrogen (H) and helium (He). The **second** period contains eight elements: lithium (Li), beryllium (Be), boron (B), carbon (C), nitrogen (N), oxygen (O), fluorine (F), and neon (Ne). The **third** period also contains eight elements beginning with sodium (Na) and ending with argon (Ar). The **fourth** period, which begins with potassium (K), and the **fifth** period, which begins with rubidium (Rb), have 18 elements each. The **sixth** period, which begins with cesium (Cs), has 32 elements. The **seventh** period, as of today, contains 32 elements, for a total of 118 elements.

Each vertical column on the periodic table contains a **group** (or family) of elements that have similar properties. A **group number** is written at the top of each vertical column (group) in the periodic table. For many years, the **representative elements** have had group numbers 1A to 8A. In the center of the periodic table is a block of elements known as the **transition elements**, which have had group numbers followed by the letter "B." A newer system assigns numbers of 1 to 18 to all of the groups going left to right across the periodic table. Because both systems are currently in use, they are both shown on the periodic table in this text and are included in our discussions of elements and group numbers. The two rows of 14 elements called the *lanthanides* and *actinides* (or the inner transition elements), which are part of Periods 6 and 7, are placed at the bottom of the periodic table to allow it to fit on a page.

- a. A period is each horizontal row in the periodic table. ____
- b. There are 8 periods. ____
- c. Each vertical column contains a group. ____
- d. A group number is written at the top of each horizontal row. ____
- e. The transition elements are in the center of the periodic table. ____
- f. There are two systems currently in use. ____
- g. *Lanthanides* and *Actinides* are placed at the top of the periodic table. ____